CHEMISTRY, Part – II

**(New Syllabus)**

**(English Version)**

**Time: 3 Hours Max.Marks:60**

SETION A

**Note: Answer all questions.** 10 ×2 =20

1. What is a buffer solution? Give one example each for acid and base buffer solutions?
2. What is the PH of a solution containing 0.63gm of HNO3 in 100 ml of solution?
3. How many oxygen’s surround phosphorus in phosphorus Pentoxide?
4. What is tailing of mercury? What is the chemical reaction taking place in it?
5. Give the composition of Nichrome?
6. What is Vulcanzation of rubber? Write the uses of polymers?
7. What is Williamson’s synthesis? Write an example?
8. What is carbylamines reaction? Give equation?
9. Define Antibiotics. Give examples?
10. Give the deficiency diseases caused by A, C, D, E, K.

SECTION B

**Note: Answer any six questions.** 6 × 4 = 24

1. Define Raoult’s law? The vapour pressure of pure water at 20° C is 17.54 mm of Hg. The Vapour pressure of the solution containing 108.24 g of non-volatile solute in 1000g of water at the same temperature in 17.354 mm of Hg. What is the molecular weight of the substance?
2. Derive Bragg’s equation?
3. State Faraday’s laws of electrolysis and explain?
4. Write any four differences between physical adsorption and chemical adsorption?
5. Define Gibbs energy?
6. Write a brief note on the extraction of sodium from sodium hydroxide?
7. Giving examples explain Werner’s theory of complex compounds?
8. What are lipids/ how are they classified?

SECTION C

**Note: Answer any two questions.** 2 × 8 = 16

1. Explain Lechatlier’s principle and apply the same to the equilibrium.

2SO2 (g) + O2 (g) **----------**> 2SO3 (g) ∆H = -189KJ

1. How is beaching powder prepared industrially? Give any its four chemical properties with equations?
2. Write three preparations, three properties and two uses of Ethyl alcohol?

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