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SRI CHAITANYA TECHNO SCHOOLS - A.P THE RIGHT MENTOR FOR IIT-JEE & OLYMPIAD

SCHOLARSHIP TEST - MODEL PAPER

IX CLASS (STATE SYLLABUS)

(Based on class VIII syllabus)

Max. Marks : 100

Read the following instructions carefully.

- 1. a) WRITE YOUR NAME AND ADDRESS completely in the space provided on Answer Sheet.
 - b) Mark the answers only on the ANSWER SHEET.
 - c) Use BLACK BALL POINT PEN for darkening the appropriate CIRCLES in the ANSWER SHEET.
 - d) Do not fold / spoil the ANSWER SHEET as it is to be evaluated by computer.
- 2. The question paper consists of 100 questions under four subject heads, Mathematics, Physics, Chemistry and English.
- 3. Number of Questions : 100 (Maths=50, Physics=15, Chemistry=15 and English=20)
- 4. All the questions are Multiple Choice type with only one correct answer and each question carries 1 mark. **NO NEGATIVE MARKS FOR ANY WRONG ANSWER**.
- 5. Return both the Question Paper and Answer Sheet to the invigilator before leaving the hall.

MATHEMATICS

1.	The difference betw of Money for 2 year	ween compound inter- ors at 4 % is Rs. 4. Th	est and simple interest en the sum is	on a certain sum	[]
	a) Rs. 3000	b) Rs. 2500	c) Rs. 4000	d) Rs.1000		
2.	A can do a piece of	f work in 8 $\frac{1}{2}$ days. A	works for 2 days and	B Joins A. They		
	work for one day. ' work is completed	The capacity of B is do by A and B with the l	ouble the capacity of A help of a boy. Then the	A. On the 4 th day t e number of days	the boy	
	can alone complete a) 30 days	ed the work is b) 25 days	c) 17 days	d) 10 days	[]
3.	In a clock the two is	hands coincide betwee	en 4 hrs and 5 hrs. Th	en the time	[]
	a) 4 hrs21 $\frac{9}{11}$ min		b) 4 hours $32\frac{4}{9}$ min	1	-	-
	c) 4 hrs $22\frac{3}{11}$ min		d) 4 hrs $20\frac{2}{11}$ min			
4.	Four carrom board touches at least two strikers is	d strikers of radius 3. o other strikers. Then	5cm are so arranged the area of empty spa	hat each striker ace between the	[]
	a) 21 sq.cms	b) 20.25 sq.cms	c) 10.5 sq.cms	d) 15 sq.cms		
5.	Fermat number is	2^{2^n} + 1. He proposed t	that $2^{2^n} + 1$ is prime for	r all 'n'		
	natural numbers.	But it fails at $n = 5$, 2^{2^2}	⁵ + 1 is not a prime.			
	Then one of the fac	ctors of 2 ^{2⁵} + 1 is			[]
	a) 231	b) 437	c) 641	d) 31		
6.	If n! (n Factorial)	represents product fir	st 'n' natural numbers	s. i.e. n!		
	$= 1 \times 2 \times 3$	n. Then number of fa	ctors of 12! is	1) 400	[]
-	a) 692	b) 792	c) 592	d) 492	r	,
7.	a) 21	b) 25	$\begin{array}{c} \text{Invisible by 2 or 3 or 5} \\ \text{c) } 26 \end{array}$	d) 24	L]
8.	The least natural r	umber which when d	ivided by 18.24 and 3	0 Leaves the		
	remainder 14,20 a	nd 26 respectively is _			[]
	a) 356	b) 354	c) 256	d) 254		
9.	If sum of all digits	in the decimal form of	f 10 ⁿ – 99 is 100, then v	value of n is	_[]
	a) 11	b) 12	c) 13	d) 14	_	_
10.	Units place digit in a) 1	the product of 22^{22} x	$x 777''' x 333^{333}$ is	d) 6	Ĺ]
11	a = 1 If $a^2 + b^2 + a^2 = D$	upper 's' and 'b' cons	c) +	$\mathbf{u} = \mathbf{u}$		
11.	$\pi a + b + c = b$	where a and b cons	ecutive positive intege	is and c – ab thei	ı r	1
	a) always an even ir	nteger	b) always an odd ir	nteger	L	1
	c) 1	<i></i>	d) irrational numbe	ſ		
12.	Ten's place digit of	f 1! + 2! + 3! +	2008! is		[]
	a) 4	b) 2	c) 1	d) None		

13.	The set builder of a soft A is	$\operatorname{set} \mathbf{A} = \{ x / -3 \le x \le 3, $	$x \in z$ }. Then total num	ber of subsets	ſ	1
	a) 64	b) 128	c) 32	d) 100	L	1
14.	Among the following	g set builder form of 'N	ull set' is		[]
	a) { $x : x \neq x$ }	b) { $x = \emptyset$ }	c) { Ø }	d) Ø		
15.	The G.C.D of two nu	umbers is 7 and the firs	st three quotients while	calculating G	C.D	
	by division Method	are 1,2 and 3. Then the	e number are	1) 01 40	[]
	a) 21, 35	b) 105, 133	c) 49, 70	d) 21, 49	_	_
16.	The descending ord a) 7 ⁹² , 8 ⁹¹ , 7 ⁸³ , 8 ¹² c) 8 ⁹¹ , 7 ⁸³ , 7 ⁹² , 8 ¹²	er of the numbers 7 ⁹² ,	8 ⁹¹ , 7 ⁸³ , 8 ¹² is b) 7 ⁹² , 7 ⁸³ , 8 ⁹¹ , 8 ¹² d) 8 ⁹¹ , 7 ⁹² , 7 ⁸³ , 8 ¹²		[]
17.	If $a^{m^n} = (a^m)^n$, then	m' in terms n is			[]
	a) n ⁿ	b) n ⁿ⁻¹	c) $n^{\frac{1}{n-1}}$	d) n ¹⁻ⁿ		
18.	Which of the followi	ng is wrong			[]
	a) $\sqrt{1.6} \times \sqrt{0.9} = 1.2$		b) $\sqrt{-9} \times \sqrt{-4} = 6$			
	c) $\sqrt{0.01} \times \sqrt{100} = 1$		d) $\sqrt{144} \times \sqrt{0.01} = 1.2$			
19.	$\mathbf{If}\left(1-\frac{1}{2}\right)\left(1-\frac{1}{3}$	$-\frac{1}{4}$)(1-	$\left(\frac{1}{n}\right) = 0.01$ then $\sqrt{n} =$		[]
	a) 0.1	b) 10	c) 1	d) 10000		
20.	The triplet (x, y, z)	satisfying the given eq	uations $x + y + xy = 5$,	$\mathbf{y} + \mathbf{z} + \mathbf{y}\mathbf{z} = 11$		
	and $z + x + z x = 7$ and $z + 3$, $3, 4$	are b) 1, 2, 3	c) 5, 4, 3	d) −1, −2, −3	[]
21.	The value of ' x ' that	t satisfying the equation	n $x - \frac{1}{x} = 3 - \frac{1}{x}$	is	[1
	a) 3	b) ()	x-3 x-3	d) No values		
22	u, s In a right angle tria	ole which sides ' r' an	d 'v' hvnotenuse 'z' th	e altitude drav	vn	
	on the hypotenuse is	'a' then	a y, nypotenuse z, m	c anniuuc uru v]]
	a) $x y = a^2$	b) $\frac{1}{x} + \frac{1}{y} = \frac{1}{a}$	c) $x^2 + y^2 = 2a^2$	d) $\frac{1}{x^2} + \frac{1}{y^2} = \frac{1}{y^2}$	$\frac{1}{a^2}$	
23.	If BC passes through given figure is, If AB	h the centre of the circl B = AC = a units	le, then the area of the s	shaded region	in the []
	a) $\frac{a^2}{2}(3-\pi)$ sq.units	b) $a^2 \left(\frac{\pi}{2} - 1\right)$ sq.	.units			
	c) $2a^2(\pi-1)$ sq.units	d) $\frac{a^2}{2} \left(\frac{\pi}{2} - 1\right)$ sq	i.units A	C		
24.	The diagonal of a sq the area of A is	uare A is $(x + y)$. The	diagonal of square B w	ith twice	[]

- the area of A is
- a) $\sqrt{2}(x+y)$ b) 2 (x+y) c) 2x+4y d) 4x+2y

25.	A square and an equ	uilateral triangle have	equal perimeters. If the	e diagonal of th	e	
	square is $6\sqrt{2}$ cm. t	hen area of the triangl	e is		[]
	a) $16\sqrt{2} \text{ cm}^2$	b) $16\sqrt{3} \text{ cm}^2$	c) $12\sqrt{2}$ cm ²	d) $12\sqrt{3} \text{ cm}^2$		
26.	A wire in shape of a wire is bent to form	n equilateral triangle e circle, the area of the c	ncloses an area of 'S' s circle is	q cm. If the san —	ne []
	a) $\frac{\pi S^2}{9}$ sq.cm	b) $\frac{3S^2}{\pi}$ sq.cm	c) $\frac{3S}{\pi}$ sq.cm	d) $\frac{3\sqrt{3.S}}{\pi}$ sq.c	m	
27.	If radius of the circl decreased by	e decreased by half, the	en the area of the circle		[]
	a) 25%	b) 75%	c) 60%	d) 80%		
28.	If $ x = \pm x$, then the	e figure formed by $ x $	+ $ y = 1$ is		[]
	a) rectangle	b) Square	c) Rhombus	d) circle		
29.	If $\left(\frac{a}{b}\right)^x = \left(\frac{b}{a}\right)^y$ then	$\frac{x}{y} + \frac{y}{x} =$			[]
	a) 4	b) 0	c) – 2	d) 2		
30.	The area of the shad	led portion in adjoinin	g figure is		[]
	a) 2000m ²		$\leftarrow^{1m} \rightarrow$	-		
	b) 90 m ²			\uparrow		
	c) 45 m ²	4	0m	Z1m		
	d) 89 m ²			¥		
31.	If a^m . $a^n = (a^m)^n$ then	n m (n–2) + n(m–2) is _	← 50		[]
	a) 1	b) – 1	c) 0	d) $\frac{1}{2}$		
32.	If $\mathbf{p} = \frac{x - y}{x + y}$, $\mathbf{q} = \frac{y - y}{y - y}$	$\frac{-z}{+z}$. $\mathbf{r} = \frac{z-x}{z+x}$ then (1 +	(p) (1 + q) (1 + r) =		[]
	a) $(1 - p) (1 - q) (1 - c) 0$	r)	b) 1 d) pqr			
33.	The triplet (x, y, z)	satisfying the given equ	$ations \frac{3xy}{x+y} = 5, \frac{2xz}{x+z} = 5$	$=3, \frac{yz}{y+z}=4 a$	re []
	a) $\frac{120}{61}, \frac{120}{11}, \frac{120}{55}$		b) $\frac{120}{31}, \frac{120}{41}, \frac{120}{23}$			
	c) 1, 2, 3		d) No solutions			
34.	The factors of x^4 + a) $(x^2 + x + 1)$ and c) $x^2 + 1$, $x^2 + 1 - x^2$	$x^{2} + 1$ are $x^{2} - x - 1$	b) $x^2 + x + 1$ and x^2 d) $x^2 + x + 1, -x^2 + 1$	$\frac{x}{x+1}$	[]
35.	The factors of (a–b) a) 3, ab, bc and ca c) 3, a, b and c	$^{3} + (b-c)^{3} + (c-a)^{3}$ are	b) 3, (a–b), (b–c) and d) No factors	(c–a)	[]

36.	The line $3x - y + 4 =$	0 doesn't passes throu	gh		[]
	a) Quadrant I	b) Quadrant II	c) Quadrant III	d) Quadrant IV	V	
37.	$Q_1 \cap Q_2 =$ a) Positive x – axis c) Negitive x – axis		b) Negitive y – axis d) null set		[]
38.	The ratio of y- coord	dinate and x – coordina	ate is	1)	[]
	a) zero	b) x – intercept	c) slope	d) y intercept		
39.	The shaded region in	n adjacent figure denot	tes 🖡 🕂	→	[]
	a) $x > 2$		(10)			
	b) $x \ge 2$		x^1 (0,0) (2,0)	× X		
	c) y > 2			\Rightarrow		
	d) $y \ge 2$		↓_ F			
40.	If \overline{BD} and \overline{CD} are t	he internal bisectors of	f the angles $ B$ and $ C $	espectively and	1	
	meet D. given that A	$= 60^{\circ}$, then $ \underline{D} = $			[]
	a) 90 ⁰	b) 120 ⁰		`		
	c) 30 ⁰	d) 100 ⁰	В	\mathbf{A}^{c}		
41.	If sum of all interior a) Nanogon	angles of regular conv b) hexagon	tex polygon is 720⁰, then c) pentagon	it is a d) octagon	.[]
42.	The ratio of interior	angle to the exterior a	ngle of a regular polygo	on is 7:2.		
	The number of sides	b) 9	c) 14	d) 16	[]
43.	The area of the circle	e in adjacent figure is	С	u) 10	ſ	1
	a) $2\pi \text{ cm}^2$	b) $\pi \text{ cm}^2$	3cm v		L	
	c) 3π cm ²	d) Data in adequate	B - 4cm - A			
44.	The line joining ince	ntre, circum centre, or	thocentre and centriod	of any		
	triangle is called as			·	[]
	a) Euler's line c) Descarte's line		b) Euclid's line d) Pythagorous line			
45.	Among the following	g measurement set, whi	ich form a triangle is		ſ	1
	a) 2, 5, 11	b) 3, 6, 10	c) 4, 5, 6	d) 5, 10, 23	-	-
46.	In the adjoing figure	$\underline{B} = 37^{0}, \ \overline{AD} = \overline{BD}$ and	d $\overline{CD} = \overline{AC}$, then $ \underline{C} =$		[]
	a) 32 ⁰	b) 48 ⁰	А			
			\bigwedge			
	c) 57 ⁰	d) 53 ⁰	B ¹³⁷⁰ B ¹¹ D ¹ C			



PHYSICS

51.	Among the following	g, the largest unit of le	ength is		[]
	a) Light year c) Parsec		b) Astronomical unit d) Nothing can be dec	ided		
52.	Wave length of sodiu a) 58960	1m is 5896 A⁰. Its valu b) 58.96	e in millimicrons is c) 5.896	d) 589.6	[]
53.	A 2 kg ball is moving The velocity with wh	g at 10ms ⁻¹ and collid nich they move combi	es with another 2 kg ball nely after collision is	l at rest.	[]
	a) 10ms^{-1}	b) 20ms^{-1}	c) $5ms^{-1}$	d) 15ms^{-1}		
54.	A solid weighs 0.08 l	kg f in air and 0.064 k	Ag f in water. The R.D of	solid is	[]
	a) 5	0) 7.5	C) 2.3	u) 10		
55.	A block of wood floa	ats in water with $\frac{1}{5}th$	of the volume above the	e surface.		
	The density of the w	rood is			[]
	a) 0.5 g / cc	b) 1 g / cc	c) 0.6 g / cc	d) 0.8 g / co	c	
56.	1°c raise of temperat a) 273	ture is equal to b) - 273	kelvin c) 272	d) 1	[]
57.	1 kg wt is equal to a) 9.8 N	b) 980 N	c) 98 N	d) 0.98 N	[]
58.	Stars twinkle due to a) refraction		b) reflection		[]

59.	The deviation product at an angle 30° is	ced in the ray, if a ra	y of light incidents on a p	plane mirroi	· ſ	1
	a) 30°	b) 60 ⁰	c) 90 ⁰	d) 120 ⁰	L	1
60.	Mark the correct sta a) The laws of reflection b) The size of virtual in c) Plane mirror always d) Plane mirror may for	tement on of light hold only for mage can be measured s form an erect image orm inverted image	or plane surfaces l by receiving it on screen		[]
61.	An object of 2 cm hig which produces a rea a) – 48 cm	gh placed at a distanc al image of 3 cm high b) – 16 cm	ce of 16 cm from a conca . The focal length of the c) – 9.6 cm	ve mirror mirror is d) – 24 cm	[]
62.	The temperature at a a) + 160 ⁰	which celcius scale re b) – 160 ⁰	ading is equal to Fahren $c) + 40^0$	heit scale is $d) - 40^0$	[]
63.	The centre of gravity a) at the centre of a lar b) at the centre of a sm c) at one of the corner d) at the point of inters	y of a regular cuboid rge face hall face s section of body diagon	is als		[]
64.	A person in a moving Then the train is said a) uniform speed c) acceleration	g train tosses a coin u l to be moving with	b) uniform velocityd) retardation		[]
65.	The number of imag kept at an angle of 72	es of an object forme 2º is	ed when two plane mirro	ors are	[]
	a) 3	b) 5	c) 6	d) 4		
		CHEM	IISTRY			
66.	Contact process is us a) Sulphuric acid c) Nitric acid	ed for the production	n of b) Sulphurous acid d) Nitrous acid		[]
67.	Magnesium oxide on turns red litmus to b a) acidic	reaction with water lue. Based on this, the	gives magnesium hydro e nature of magnesium o	xide which oxide is d) amphote	[pric]
68.	The metallic oxides of at elevated temperat CuO + H ₂ $\uparrow \rightarrow$ Cu + a) Copper is reduced c) Copper is neutralise	can be converted to m ures. This process is s + H ₂ O. In this reactio	netals by passing pure hy shown in the following en n b) Copper is oxidised d) Hydrogen is reduce	d rogen quation	[]
69.	The amount of oxyge the reaction 2Mg + C a) 160g	en required to combin $D_2 \rightarrow 2 \text{ MgO is}$ b) 1.6g	ne with 2.4grams of mag	g nesium in d) 32g	[]
70.	A particular nitrogen compound. Is anothe Here, the law followe	n oxide containing 0.8 er oxide possible with ed is	8775g of nitrogen per gra 0.2925g of nitrogen per	am of the gram.	[]
	a) No, Law of conserv	vation of mass	b) Yes, Law of multiple	e proportions		
	cal proportion	ns				

71. Among the following, the incorrect statement is

- a) Action of heat on NH₄Cl is a physical change
- b) NH₄Cl is a sublimable compound
- c) Action of heat on NH₄Cl is a chemical change
- d) NH₄Cl on heating gives white dense fumes.

72. Read the following reactions, identify A, B, C and D [] heat heat

	_{a)} ZnO		ZnO		b) PbO
	a (A)	cool (B)		^(C) (C)
	Α	В	C	D	
a)	White	yellow	yellow	brown	
b)	White	yellow	brown	yellow	
c)	White	brown	yellow	white	
d)	White	brown	yellow	brown	

Statement 1 : Formula of carbonate ion is CO_3^{2-} 73. Statement 2 : Formula of sulphite ion is S^{-2}

Statement 3 : Formula of dichromate ion is $\operatorname{Cr}_2 O_7^{-2}$

Option	Statement 1	Statement 2	Statement 3
a)	Т	Т	Т
b)	Т	F	Т
c)	Т	F	F
d)	F	Т	Т

- 74. Electrolysis of water is shown in the figure. Identify A, B, C and D
- B D A С 0, a) Anode cathode H₂ O_2 b) Cathode anode H₂ Anode cathode 0, H₂ c) d) Cathode Anode Ο, H₂
- 75. 1.375g of pure cupric oxide is reduced by heating in a current of pure and dry hydrogen and the mass of copper that remained is 1.0980g. In another experiment, 1.179g of pure copper is dissolved in pure HNO₃ and the resulting copper nitrate converted into copper oxide by ignition. The mass of copper oxide formed is 1.476g. The above data illustrates a) Law of conservation of mass b) Law of multiple proportions
 - c) Law of constant composition
- Read the following reactions and choose the correct option 76. 1. $NH_3 + HCl \rightarrow NH_4Cl$ 2. 2Pb $(NO_3)_2 \rightarrow 2PbO + O_2 \uparrow + 4 NO_2 \uparrow$

3.
$$Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$$

4. AgNO, $+ HCl \rightarrow AgCl + HNO$,

a	b	с	d
1	2	4	3
2	3	3	2
3	1	1	1
4	4	2	4
	a 1 2 3 4	a b 1 2 2 3 3 1 4 4	a b c 1 2 4 2 3 3 3 1 1 4 4 2

Γ

]

[]

]

Γ

+B $A_{\mathcal{A}}$

d) Law of reciprocal proportions

PbO

(D)

cool

]

]

Γ

[

77.	Mat	ch the fo	llowing	5				[]
	Set A		Set B		Set C				
1.	Sodium		A) Cupru	m	a) Ag				
2.	Copper		B) Argent	tum	b) Cu				
3.	Silve	er		C) Aurum	1	c) Pb			
4.	Gold	1		D) Plumb	oum	d) Na			
5.	Leac	1		E) Natriu	m	e) Au			
	a)	1	2	3	4	5			
		E	A	В	С	D			
		d	b	a	e	С			
	b)	1	2	3	4	5			
		A	В	С	E	D			
		d	b	a	С	e			
	c)	1	2	3	4	5			
		E	A	В	D	С			
		d	b	e	a	С			
	d)	1	2	3	4	5			
		E	A	С	В	D			
		d	b	С	a	e			
78.	I hay	ve 100gn	n of Ca	CO ₃ . If I he	at it strongl	y for some time in	an open dish,	-	-
	cool	it and w	eigh it a	again, its w	eight would	be	1) 40	Ĺ]
	a) 95	ogm		b) 56gm		c) 44gm	d) 40gm		
79.	2KN	$4 \ln O_4 + 3$	H ₂ SO ₄	$+5H_2S\uparrow =$	$\rightarrow \mathbf{K}_{2}\mathbf{SO}_{4} + 2$	$MnSO_4 + 8H_2O +$	5 S		
79.	2KM In th	$InO_4 + 3$ ne above	H ₂ SO ₄ reactio	+ $5H_2S\uparrow =$ n oxidising		$MnSO_4 + 8H_2O +$	58	[]
79.	2KN In th a) K	$4 nO_4 + 3$ ne above MnO_4	⁶ H ₂ SO ₄ reactio	+ $5H_2S \uparrow =$ n oxidising b) H_2SO_4	$ ightarrow K_2 SO_4 + 2$ agent is	$MnSO_4 + 8H_2O +$ c) H ₂ S	5S d) H ₂ O	[]
79. 80.	2KM In th a) K	$InO_4 + 3$ ne above MnO_4	³ H ₂ SO ₄ reactio	+ $5H_2S\uparrow =$ n oxidising b) H_2SO_4	$ ightarrow K_2SO_4 + 2$ agent is	$MnSO_4 + 8H_2O +$ c) H ₂ S	5S d) H ₂ O	[]
79. 80.	2KM In th a) K	$InO_4 + 3$ ne above MnO_4 Dxide	³ H ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4	$\Rightarrow K_2 SO_4 + 2$ agent is e of oxide	$MnSO_4 + 8H_2O +$ c) H ₂ S	5S d) H ₂ O	[]]
79. 80.	2KM In th a) K 0 1) P	$\frac{\text{InO}_4 + 3}{\text{ne above}}$ $\frac{\text{MnO}_4}{\text{Dxide}}$	³ H ₂ SO ₄ reactio	+ $5H_2S\uparrow -$ n oxidising b) H_2SO_4	$\Rightarrow K_2 SO_4 + 2$ agent is $e \text{ of oxide}$	$\mathbf{MnSO}_{4} + \mathbf{8H}_{2}\mathbf{O} +$ c) $\mathbf{H}_{2}\mathbf{S}$	5S d) H ₂ O	[[]
79. 80.	2KM In th a) K 0 1) P 2) C	$\frac{\text{InO}_4 + 3}{\text{ne above}}$ $\frac{\text{MnO}_4}{\text{Oxide}}$ $\frac{2\text{O}_5}{\text{O}_2}$	² H ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4 Nature	→ K ₂ SO ₄ + 2 agent is e of oxide	$\mathbf{MnSO}_{4} + \mathbf{8H}_{2}\mathbf{O} +$ c) $\mathbf{H}_{2}\mathbf{S}$	5 S d) H ₂ O]]
79. 80.	2KM In th a) K 0 1) P 2) C 3) M	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{Dxide}}$ $\frac{\text{Dxide}}{\text{O}_{5}}$ $\frac{\text{O}_{7}}{\text{IgO}}$	² H ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4	→ K ₂ SO ₄ + 2 agent is e of oxide	$\mathbf{MnSO}_4 + \mathbf{8H}_2\mathbf{O} + \mathbf{C} + \mathbf{H}_2\mathbf{S}$	5S d) H ₂ O]]
79. 80.	2KN In th a) K (0) 1) P 2) C 3) M 4) S($\frac{\text{InO}_4 + 3}{\text{ne above}}$ $\frac{\text{MnO}_4}{\text{Oxide}}$ $\frac{\text{Oxide}}{\text{O}_2}$ $\frac{\text{O}_2}{\text{O}_2}$	PH ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4	→ K ₂ SO ₄ + 2 agent is e of oxide	$MnSO_4 + 8H_2O +$ c) H ₂ S	5S d) H ₂ O]]
79. 80.	2KN In th a) K (0) 1) P 2) C 3) M 4) S(0)	$\frac{\text{InO}_4 + 3}{\text{ne above}}$ $\frac{\text{MnO}_4}{\text{Oxide}}$ $\frac{\text{Oxide}}{\text{O}_2}$ $\frac{\text{O}_2}{\text{IgO}}$ 1	PH ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4 Nature	$\Rightarrow K_2 SO_4 + 2$ agent is $e \text{ of oxide}$	$\mathbf{MnSO}_{4} + \mathbf{8H}_{2}\mathbf{O} + \mathbf{C} + \mathbf{H}_{2}\mathbf{S}$	5 S d) H ₂ O]]
79.	2KM In th a) K 0 1) P. 2) C 3) M 4) S0	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{Oxide}}$ $\frac{\text{Oxide}}{\text{O}_{5}}$ $\frac{\text{O}_{5}}{\text{O}_{2}}$ $\frac{1}{1}$ $\frac{1}{1}$	PH ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4 Nature	$\Rightarrow K_2 SO_4 + 2$ agent is $e \text{ of oxide}$ 3 Basis	$MnSO_4 + 8H_2O +$ c) H ₂ S	5 S d) H ₂ O	[]
79.	2KN In th a) K (1) P 2) C 3) M (4) S (2) (3) M (4) S (3)	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{MnO}_{4}}$ $\frac{\text{Dxide}}{\text{O}_{5}}$ $\frac{\text{O}_{7}}{\text{O}_{2}}$ $\frac{1}{\text{Acidic}}$	PH ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow n$ oxidising b) H_2SO_4 Nature	$\Rightarrow K_2 SO_4 + 2$ agent is $e \text{ of oxide}$ 3 Basic Device	$MnSO_4 + 8H_2O +$ c) H ₂ S	5 S d) H ₂ O]]
79.	2KM In th a) K (0) (1) P, (2) C (3) M (4) S(a) b)	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{Oxide}}$ $\frac{\text{Oxide}}{\text{O}_{2}}$ $\frac{\text{O}_{2}}{1}$ $\frac{1}{\text{Acidic}}$ $\frac{1}{1}$	PH ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow$ n oxidising b) H_2SO_4 Nature 2 Basic Acidic	$\Rightarrow K_2SO_4 + 2$ agent is $e \text{ of oxide}$ 3 Basic Basic	$MnSO_4 + 8H_2O +$ c) H ₂ S 4 Acidic Acidic	5 S d) H ₂ O]]
79.	2KN In th a) K (1) P (2) C (3) M (4) S (4) S (4) S (5) (5) (5) (5) (5) (5) (5) (5) (5) (5)	$InO_4 + 3$ In above MnO_4 Dxide O_2	PH ₂ SO ₄ reactio	+ $5H_2S \uparrow \rightarrow$ n oxidising b) H_2SO_4 Nature 2 Basic Acidic Acidic	$\Rightarrow K_2SO_4 + 2$ agent is $e \text{ of oxide}$ $g \text{ of oxide}$ 3 Basic Basic Acidic	$MnSO_4 + 8H_2O +$ c) H_2S 4 Acidic Acidic Basic	5 S d) H ₂ O]]
79.	2KN In th a) K (0) (1) P (2) C (3) M (4) S((3) (4) S((3) (4) S((3) (4) S((3) (4) S((3) (4) S((3) (4) S((3) (4) S((3)) (4) S((3)) (4) S((3)) (4) S((3)) (4) S((3)) (4) S((3)) (4) S((3)) (4)) (4)) (4)) (4)) (4)) (4)) (4)	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{MnO}_{4}}$ $\frac{\text{Dxide}}{\text{O}_{2}}$ $\frac{\text{O}_{2}}{\text{O}_{2}}$ $\frac{1}{\text{Acidic}}$ $\frac{\text{Acidic}}{\text{Acidic}}$ $\frac{\text{Acidic}}{\text{Acidic}}$	² H ₂ SO ₄ reactio	+ $5H_2S$ ↑ → n oxidising b) H_2SO_4 Nature 2 Basic Acidic Acidic Basic	\Rightarrow K ₂ SO ₄ + 2 agent is e of oxide \approx Basic Basic Acidic Acidic	$\mathbf{MnSO}_{4} + \mathbf{8H}_{2}\mathbf{O} +$ c) $\mathbf{H}_{2}\mathbf{S}$ $\frac{4}{\mathbf{Acidic}}$ \mathbf{Acidic} \mathbf{Basic} \mathbf{Basic}	5S d) H ₂ O]]
79.	2KN In th a) K (1) P (2) C (3) M (4) S(a) b) c) d)	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{MnO}_{4}}$ $\frac{\text{Dxide}}{\text{O}_{2}}$ $\frac{\text{O}_{2}}{\text{O}_{2}}$ $\frac{1}{\text{Acidic}}$ $\frac{\text{Acidic}}{\text{Acidic}}$ $\frac{\text{Acidic}}{\text{Acidic}}$	PH ₂ SO ₄ reactio	+ $5H_2S$ ↑ → n oxidising b) H_2SO_4 Nature 2 Basic Acidic Acidic Basic	→ K ₂ SO ₄ + 2 agent is 2 of oxide 3 Basic Basic Acidic Acidic ENGI	$\mathbf{MnSO}_{4} + \mathbf{8H}_{2}\mathbf{O} +$ c) $\mathbf{H}_{2}\mathbf{S}$ $\begin{bmatrix} 4 \\ Acidic \\ Acidic \\ Basic \\ Basic \\ Basic \\ \mathbf{LISH} \end{bmatrix}$	5S d) H ₂ O]]
79. 80.	2KN In th a) K (1) P (2) C (3) M (4) S (3) (4) S (3) (4) S (3) (4) S (3) (4) S (3) (4) S (3) (4) S (3) (4) S (3) (4) S (4) S ($\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{MnO}_{4}}$ $\frac{\text{Dxide}}{\text{O}_{2}}$ $\frac{\text{O}_{2}}{\text{O}_{2}}$ $\frac{1}{\text{Acidic}}$ $\frac{\text{Acidic}}{\text{Acidic}}$ $\frac{\text{Acidic}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$ $\frac{\text{Oxide}}{\text{Acidic}}$	PH ₂ SO ₄ reactio	+ $5H_2S$ ↑ → n oxidising b) H_2SO_4 Nature 2 Basic Acidic Acidic Basic	K₂SO₄ + 2 agent is agent is agent is 3 Basic Basic Acidic Acidic ENGI of the verb	MnSO ₄ + 8H ₂ O + c) H ₂ S 4 Acidic Acidic Basic Basic LISH given.	5S d) H ₂ O]]
79. 80. I.	2KN In th a) K (1) P (2) C (3) M (4) S(1) (4) S(1) (5) (6) (7) (7) (7) (7) (7) (7) (7) (7) (7) (7	$\frac{\text{InO}_{4} + 3}{\text{ne above}}$ $\frac{\text{MnO}_{4}}{\text{MnO}_{4}}$ $\frac{\text{Dxide}}{\text{O}_{2}}$ $\frac{\text{O}_{2}}{\text{O}_{2}}$ $\frac{1}{\text{Acidic}}$ Acidic Acidic Acidic Acidic One of	PH ₂ SO ₄ reactio	+ $5H_2S$ ↑ → n oxidising b) H_2SO_4 Nature 2 Basic Acidic Acidic Basic table form ends	K₂SO₄ + 2 agent is agent is	MnSO ₄ + 8H ₂ O + c) H ₂ S 4 Acidic Acidic Basic Basic UISH given. me if I had aske	5S d) H ₂ O]]
79. 80. I.	2KN In th a) K (1) P, 2) C (3) M (4) S(a) b) c) d) Cho 81.	$\frac{\text{InO}_4 + 3}{\text{ne above}}$ $\frac{\text{MnO}_4}{\text{MnO}_4}$ $\frac{\text{Dxide}}{2O_5}$ $\frac{O_7}{O_2}$ $\frac{1}{Acidic}$ $Acidic$ $Acid$ $Acidic$ $Acidic$ $Acidic$ $Acidic$ $Acid$ $Acidic$	nost sui my frie	+ $5H_2S$ ↑ → n oxidising b) H_2SO_4 Nature 2 Basic Acidic Acidic Basic table form ends ped	→ K ₂ SO ₄ + 2 agent is e of oxide 3 Basic Basic Acidic Acidic ENGI of the verb	MnSO ₄ + 8H ₂ O + c) H ₂ S 4 Acidic Acidic Basic Basic ElISH given. _ me if I had aske b) have helped	5S d) H ₂ O]]

	82.	I didn't see				[]		
		a) where are ye	ou ?	b) where were you ?					
		c) where you v	were	d) where you have	been				
	83.	What a beaut	iful horse			[]		
		a) it is ?	b) is it ?	c) it is !	d) is it !				
II.	Cho	ose the correct	alternative.						
	84.	Though the ra	ainfall was adequat	e this year, the mango tre	ees still				
		a high yield				[]		
		a) had not proc	luced	b) did not produce					
		c) would not p	roduce	d) was not produced	d				
	85.	It is earth's g	ravity which	people their v	veight	[]		
		a) given	b) gave	c) giving	d) gives				
	86.		president, a ca	ndidate must win a majo	rity of votes.	[]		
		a) Elected a		b) To be elected as					
		c) Electing as		d) To be elected a					
	87.	Total weight o	of all the ants in the	world is much greater th	nan	_[]		
		a) all human b	eings	b) to all human bein	igs				
		c) that of all hu	ıman beings	d) of all human beir	igs				
	88.		, I can get so ma	ny books to read.		[]		
		a) My father is	a librarian	b) Being my father a	a librarian				
		c) My father is	a librarian but	d) My father being a	a librarian				
III.	The	re is an error ir	n each of the senten	ces given below. Choose	the correct				
	alter	rnative and wri	te its option in the b	orackets.					
	89.	He does not s	moke, <u>nor he drink</u>	<u>s</u>		[]		
		a) nor he does	drink	b) nor does he drink	X				
		b) neither he d	oes drink	d) no change					
	90.	The notes <u>is w</u>	vriting by me now			[]		
		a) has been wr	iting	b) was being written	1				
		c) is being writ	tten	d) has written					
IV.	Identify the correct parts of speech of the following.								
	91.	I complained	about my stomach	ache		[]		
		a) verb		b) adjective					
		c) conjunction		d) preposition					
	92.	They like him	because he <u>is hone</u>	st		[]		
		a) preposition		b) conjunction					
		c) verb		d) pronoun					
V.	Cho	ose the correct	meaning of the wor	rd underlined					
	93.	We were <u>fasci</u>	inated by the speech	n of our friend		[]		
		a) amused	b) satisfied	c) disgusted	d) attracted				
VI.	Cho	ose the wrongly	y - spelt word from	the words given below		[]		
	94.	a) arguement	b) divide	c) opinion	d) precious				

VII.	VII. Choose the best meaning for the idiom given below						
	95.	A snake in the gamma a) a very poisone c) an unrecognis	grass ous snake able enemy	b) a secret agent d) not a reliable p	person	[]
VIII	. Cho	ose the correct se	entence of the follo	owing			
	 96. His pocket has been picked a) Some one have picked his pocket b) Some one picked his pocket c) Picked has been his pocket d) Some one has picked his pocket 					[]
	97.	The teacher punished the boys who had not done their home work.[]a) The boys who had not done their home work had been punished by their teacher.b) The boys were punished by their teacher who had not done their home work.c) The boys who had not done their home work were punished by the teacher.d) The boys who had not done their home work were being punished by the teacher.					
IX.	. Identify the degree of comparison						
	98.	Pen is mightier a) superlative	than sword b) comparative	c) positive	d) compour	[nd]
X.	Select the correct word or phrase to complete a grammatical sentence.						
	99.	They enjoyed during holidays.				[]
		a) in	b) themself	c) themselves	d) None		
XI.	Complete the following proverb choosing the best from the given below						
	100.	 Birds of same feather a) flock together c) gather food together 		b) build their nes d) hunt their prey	b) build their nest together d) hunt their prey together]
